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Trends in the Periodic Table

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Periodic Trends - 6. Locate the numbers in Model 1 that represent the ionization energy. The ionization energy is the amount of energy needed to remove an electron from an atom. a. Using your knowledge of Coulombic attraction, explain why ionization—removing an

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Across the periodic table, sizes of atoms show the following trends, with many irregularities: Size increases down a group. The outermost electrons are in successively more extensive orbitals as n increases. Size decreases across a period. Electrons are added to the same shell and do not shield one

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Periodic trends are specific patterns that are present in the periodic table that illustrate different aspects of a certain element, including its size and its electronic properties. Major periodic trends include: electronegativity, ionization energy, electron affinity, atomic radius, melting point, and metallic character .

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The trends in these properties as you go across periods and down groups is the subject of this Chem POGIL Success Criteria: • Understand the meaning of atomic radius, reactivity, electronegativity, ionization energy • Recognize trends in atomic radius, reactivity, ionization energy and electronegativity as you go across periods and down groups.

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