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Calorimetry, Specific
Heat, and
Calculations Example
Question #1 :
Calorimetry, Specific
Heat, And

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Calculations The following is a list of specific heat capacities for a few metals.

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Measure the
temperature of the
water in the
calorimeter and
record it (to $\pm 0.10^\circ\text{C}$)
as the initial
temperature of cold

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water. Place 100.0 mL water in a clean, dry beaker. Heat the water in the beaker until it is at least 300C warmer than the cold water.

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Activity/Student Guide

Measure the temperature of the water in the calorimeter and record it (to IOC) as the initial temperature of cold water. Place 100.0 mL water in a clean, dry beaker. Heat the water in the beaker until it is at least 30°C warmer than the cold water.

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the size of an atom
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of Calorimetry

The law itself says nothing about standard enthalpies of formation. Any way that we can figure out the change in heat between the products and the reactants will work just as well. And that's where calorimetry comes in. Calorimetry is the science of measuring

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the change in heat associated with a chemical reaction.

Calorimetry and enthalpy introduction (video)

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Calorimetry: - the science of measuring heat flow. - heat is defined as thermal energy flowing from an object at a higher

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temperature to one at a lower temperature. Exothermic reactions release energy or heat to increase temperature of the surrounding.

CHM2046L
Calorimeter Lab by
Nupur Godbole
PART I:
Fundamentals of
Calorimetry

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Differences Summary
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the measurement of
heat flow Heat energy
flows from a
substance that has a
higher temperature to
a substance that has
a lower temperature
The heat will continue
to flow until both
substances reach the
same temperature
known as the final

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