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Solving the chemical  
equilibria problem boils  
down to finding the extents  
of reaction and temperature  
that equate the equilibrium  
coefficient calculated both

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ways. Consider the reaction  
 $A + B \rightleftharpoons C + D$  (12)  
with initial mole fractions  
of each species and moles of  
each species given by

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Chemical equilibrium, a condition in the course of a reversible chemical reaction in which no net change in the amounts of reactants and products occurs. A reversible chemical reaction is one in which the products, as soon as they are formed, react to produce the original reactants. At equilibrium, the two opposing reactions go on at equal rates, or velocities, hence there is no net change in the amounts of substances involved.

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Given the equation,  $N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$ ,  
Find  $Q$  and determine which direction the reaction will

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shift in order to reach the state of chemical equilibrium. Solution: Given,  $[N_2] = 0.04M$ ,  $[H_2] = 0.09M$ , and  $K = 0.040$

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physicochemical environment  
of a metabolite.

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Chemical Equilibrium

Utkstair Solving the

chemical equilibria problem

boils down to finding the

extents of reaction and

temperature that equate the

equilibrium coefficient

calculated both ways.

Consider the reaction ?A A +

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?B B ? ?C C +?DD(12) with  
initial mole fractions of  
each species and moles of  
each species given by  
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Third, we have chemical  
equilibrium, equating the  
chemical species of each  
comonent in the two phases,

$\mu_i^L = \mu_i^V$  for  $i = 1$  to  $N_c$ ,

(3)

~~The Thermodynamics of Vapor-~~

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This chemistry video

tutorial provides a basic



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introduction into how to solve chemical equilibrium problems. It explains how to calculate the equilibrium co...

~~How To Calculate The  
Equilibrium Constant K  
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Chemical equilibrium is the state of a reversible reaction where the rate of the forward reaction equals the rate of the reverse reaction. While a reaction is in equilibrium the concentration of the reactants and products are constant.

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